

CATALYTIC REFORMING CATALYST ACTIVATION

This invention relates to the regeneration and activation of reforming catalyst and the use of such activated catalyst in the reforming of hydrocarbons.

BACKGROUND OF THE INVENTION

Catalytic reforming, or hydroforming, is a well established industrial process employed by the petroleum industry for improving the octane quality of naphthas or straight run gasolines. In reforming, a multi-functional catalyst is employed which typically contains a metal hydrogenation-dehydrogenation (hydrogen transfer) component or components, substantially atomically dispersed upon the surface of a porous inorganic oxide support, notably alumina.

In recent years, platinum has been widely commercially used as the metallic hydrogen transfer component of reforming catalysts, and

platinum-on-alumina catalysts have been commercially employed in refineries. Also, additional metallic components, such as rhenium, iridium, ruthenium, tin, palladium, germanium and the like, have been added to platinum as promoters to further improve the activity, selectivity, or both, of the basic platinum catalyst.

In a conventional reforming process, a series of reactors constitute the heart of the reforming unit. Each reforming reactor is generally provided with a fixed bed or beds of the catalyst which receive upflow or downflow feed. Each reactor is provided with a heater because the reactions which take place therein are endothermic. In a conventional reforming process, a naphtha feed with hydrogen or hydrogen recycle gas is passed through a preheat furnace, then downward through a reactor, and then in sequence through subsequent interstage heaters and reactors of the series. The product of the last reactor is separated into a liquid fraction and a vaporous effluent. The vaporous effluent, a gas rich in hydrogen, is used as hydrogen recycle gas in the reforming process.

During operation, the activity of the reforming catalyst gradually declines due to the build-up of coke, and the temperature of the process is gradually raised to compensate for the activity loss caused by the coke deposits. Eventually, economics dictate the necessity of regenerating the catalyst.

The initial phase of catalyst regeneration is accomplished by burning the coke off the catalyst under controlled conditions. Catalyst regeneration is then completed through a sequence of activation steps wherein the agglomerated metal hydrogenation-dehydrogenation components are atomically redispersed. Such activation generally is achieved by treating the catalyst with hydrogen to effect reduction of the platinum oxide, and such other oxides as may be present in the catalyst system, followed by a chloride treatment of the reduced catalyst system prior to placing it back into use.

In addition to the activation process required when working with a catalyst which has been subjected to regeneration by burning off coke from the catalyst, there is generally carried out an activation treatment of the initially charged catalyst to the reactor prior to the introduction of hydrocarbon feed to the system.

In both the activation of fresh catalyst as well as the activation of a regenerated catalyst there have been numerous efforts to achieve a catalyst system whereby the catalyst will have increased activity, provide a product having increased octane values and avoid the undesirable cracking of product.

It is an object of the present invention to provide a process for the regeneration and activation of a platinum reforming catalyst.

Another object of the invention is to provide an improved system for the activation of a platinum reforming catalyst whereby the catalyst activity is increased.

5 A still further object of this invention is to provide an improved process for the activation of a platinum reforming catalyst which, when the activated catalyst is used in the reforming of a reformer feed, the cracking of such feed to undesired byproducts is reduced.

10 Yet another object of this invention is to provide a process for the activation of a platinum reforming catalyst which, when the activated catalyst is used in the reforming of a reformer feed, the cracking of such feed is decreased while the octane number of the reformer product is increased.

Other aspects, objects and the several advantages of the invention will be apparent from the following specification and appended claims.

SUMMARY OF THE INVENTION

15 In accordance with the present invention, an improved reforming catalyst is obtained when the catalyst, during reduction with hydrogen, is simultaneously contacted with hydrogen and a nonmetallic chlorine-containing compound in a reactor of a series of multiple reactors, and thereafter a hydrogen purge is maintained for a sufficient amount of time to expose the

reforming catalyst to about 100 to about 50,000 cubic feet of hydrogen per cubic foot of catalyst prior to bringing the system to reforming conditions.

Thus, in accordance with one aspect of the present invention, there is provided a process for the activation of a platinum reforming catalyst contained in a multiple reaction zone system, said process comprises: (a) reducing the catalyst with hydrogen; (b) simultaneously with step (a) contacting the catalyst with a nonmetallic chlorine-containing compound by introducing the nonmetallic chlorine-containing compound into a reaction zone of the multiple reaction zone system under conditions to effect decomposition of the nonmetallic chlorine-containing compound; and (c) thereafter purging the resulting chlorine-treated catalyst with hydrogen for a period of time necessary to remove excess chlorine from the catalyst prior to the use of said chlorine-treated catalyst in a reforming process.

In accordance with another aspect of the present invention, there is provided an improved process for the regeneration of a platinum reforming catalyst contained in a multiple reaction zone system, said process comprises: (a) purging the multiple reaction zone system with nitrogen; (b) subjecting the deactivated catalyst to an oxidative burning off at a temperature and for a time sufficient to remove substantially all carbonaceous deposits thereon; (c) subjecting the substantially-carbon-free catalyst to an oxygen treatment for a

period of time sufficient to effect the oxidation of metals contained in the substantially-carbon-free catalyst; (d) purging the resulting oxidized catalyst of molecular oxygen; (e) cooling the resulting purged catalyst; (f) reducing the dried catalyst by contacting with hydrogen which is introduced into a reaction zone of the multiple reaction zone system; (g) simultaneously with step (f) contacting the catalyst with a nonmetallic chlorine-containing compound by introducing the chlorine-containing compound into a reactor of the multiple reaction zone system; and (h) thereafter purging the resulting catalyst systems of steps (f) and (g) with hydrogen for a period of time necessary to remove excess chlorine from the catalyst prior to start up of the multiple reaction zone reforming system.

DETAILED DESCRIPTION OF THE INVENTION

In carrying out the process of the present invention there is employed a treatment of the catalyst beds of a reforming system which employs a series of reactors, generally three or four, which can contain varying catalyst compositions in each of the reactors. The individual reactors of the series can contain a platinum-alumina catalyst system either alone or in combination with an additional metallic compound such as rhenium, iridium, ruthenium, tin, palladium, germanium, and the like. It is presently preferred, in carrying out

the regeneration and activation processes of this invention that a platinum-rhenium-alumina catalyst system be utilized in each reactor of the series.

The present invention is based upon the discovery that in carrying out the generally practiced procedure for regeneration of reforming catalyst contained in a series of reactors, wherein the deactivated catalyst is subjected to oxidation to remove carbonaceous deposits and thereafter activated through the addition of a chlorine-containing compound and hydrogen to the reactor system, or that fresh catalyst is to be activated with hydrogen and a chlorine-containing compound, that by introducing a nonmetallic chlorine-containing compound simultaneously with hydrogen to at least one of the reactors, preferably to only the first of the reactors of the series and thereafter purging the resulting activated catalyst system with from about 100 to about 50,000 cubic feet of hydrogen per cubic feet of catalyst there is achieved a suitably activated catalyst which, when placed in service in the reforming process, provides a system in which undesirable cracking of the feedstock decreases while octane number of the reforming process product and catalyst activity increases.

In accordance with this invention, a reforming catalyst containing a platinum metal is activated by simultaneously contacting the catalyst with

hydrogen and a nonmetallic chlorine-containing compound, and thereafter purging the thus-treated catalyst with hydrogen.

An essential aspect of this invention is that the chloride treatment of the catalyst must be performed simultaneously with hydrogen reduction of the catalyst. Chloride treatment is carried out by adding a chlorine-containing compound to a reforming reactor while simultaneously reducing the catalyst within the reactors with hydrogen. The chlorine-containing compound employed in the chloride treatment is preferably a nonmetallic chloride.

Nonmetallic chlorine-containing compounds are preferred; because, contacting the catalyst with metallic chlorine-containing compounds tends to result in the build-up on the catalyst of undesirable metals, which lead to reduced activity, reduced selectivity, or both. More preferably the chlorine-containing compound is a nonmetallic organic chloride. Organic chlorides are preferred because they tend to decompose better under activation conditions in the reformer. Preferred nonmetallic organic chlorides include, for example, tetrachloroethylene, hexachloroethane, carbon tetrachloride, 1-chlorobutane, 1-chloro-2-methyl propane, 2-chloro-2-methyl propane, tertiary butyl chloride, propylene dichloride, perchloroethylene, and mixtures of two or more thereof. The presently most preferred non-metallic chloride is perchloroethylene.

Generally, the quantity of chlorine-containing compound employed during the chloride treatment must be sufficient to provide in the catalyst system from about 0.05 to about 0.3 pounds of chlorine per pound of catalyst, preferably from about 0.1 to about 0.2 pounds of chlorine per pound of catalyst. The temperature employed during chloride treatment must be sufficient so as to effect decomposition of the chlorine-containing compound. The chloride treatment can be performed at a temperature of from about 500°F to about 1,500°F, preferably from about 700°F to about 1,200°F, and most preferably from about or 900°F to or about 940°F, and a pressure in the range of about 0 to about 600 psig, preferably about 50 to about 300 psig.

As mentioned above, the chloride treatment must occur simultaneously with, and under the same conditions as, reduction of the catalyst with hydrogen. The quantity of hydrogen within the system during chloride treatment must be equal to the stoichiometric amount required to form hydrogen chloride with the chlorine obtained from the chlorine-containing compound. At the same time there must be a substantial absence in the system of free oxygen or compounds that decompose at the treating conditions to produce free oxygen.

Following chloride treatment and reduction with hydrogen, it is essential that the system is purged with an amount of hydrogen necessary to

remove excess chlorine from the catalyst. The amount of hydrogen employed during the purging process can be from about 100 to about 50,000, preferably from about 500 to about 30,000, and most preferably 1,000 to 10,000 cubic feet of hydrogen per cubic foot of catalyst. The hydrogen purge can be conducted at a temperature from about 500°F to about 1,500°F, preferably about 800°F to 1,100°F, and most preferably from 900°F to 940°F, and a pressure in the range of about 0 to about 600 psig, preferably about 50 to about 300 psig.

If the hydrogen purge is conducted in a manner which exposes the catalyst to less than about 100 cubic feet of hydrogen per cubic foot of catalyst, the resulting catalyst will not meet the objects of this invention due to the existence of excessive chlorine on the catalyst. If the hydrogen purge is conducted in a manner which exposes the catalyst to more than about 50,000 cubic feet of hydrogen per cubic foot of catalyst, the resulting catalyst will not meet the objects of this invention because the catalyst will not contain sufficient chlorine.

While the above-described technique is applicable to the activation of substantially carbon-free fresh catalyst, it is particularly advantageous to incorporate such technique in a method for reactivating catalysts which have become deactivated through employment in a reforming process (i.e., a combined regeneration-reactivation procedure).

In accordance with this embodiment, the invention is directed to a method for regenerating and activating a reforming catalyst containing platinum metals that has become deactivated through a series of reforming-regeneration cycles. This method comprises first purging the deactivated catalyst with an inert gas, such as nitrogen. Then subjecting the deactivated catalyst to an oxidative burn off at a temperature and period of time sufficient to remove substantially all carbonaceous deposits therefrom. Quite commonly this oxidative burn-off is accomplished in two phases, the first of which is principally a carbon burn-off phase, while the second phase can be termed an afterburn.

In accordance with this technique, the carbon burn-off of the inventive process can be effected by contacting the carbon-containing catalyst with an oxygen-containing gas, preferably a molecular oxygen-containing gas, having a comparatively low oxygen content which is less than about 1 percent by volume, more preferably, between 0.2 and 0.8 percent by volume. The temperature employed in the carbon burn-off technique ranges from a minimum temperature necessary to effect oxygen-containing of the carbon in the presence of the oxygen-containing gas up to a maximum temperature of about 1,200°F, preferably from about 300°F to about 850°F.

In the afterburn phase, generally a gas of a comparatively higher oxygen content is employed, e.g. up to about 2 percent by volume of oxygen.

In this afterburn technique the maximum temperature employed is about 1,300°F, preferably the temperature should be from about 400°F to about 950°F. The temperature in both the carbon burn-off and afterburn phases can readily be controlled by adjusting of the oxygen content of the oxygen-containing gas.

The catalyst resulting from the oxidative burn off should be substantially carbon free, with from 80 percent upwardly to 100 percent of the carbon originally on the catalyst being burned off. The substantially carbon free catalyst will thus generally have less than about one weight percent carbon but, more particularly, such substantially carbon free catalyst will contain carbon in the range of from about 0.01 weight percent to about 0.75 weight percent, but, preferably, from 0.01 to 0.25 weight percent. Most preferably, the carbon content will be less than 0.15 weight percent, thus, from 0.01 to 0.15 weight percent. Generally, the oxidative burn off should be performed from about 4 to about 36 hours, preferably from 8 to 18 hours.

After the oxidative burn off, the substantially-carbon-free catalyst is subjected to an oxygen treatment with a gas containing at least about 5 percent by volume of molecular oxygen at a temperature in the range of from

about 800°F to about 1,150°F, preferably from about 900°F to about 940°F. In this oxygen treatment step of the invention, any suitable oxygen-containing gas can be employed including, for example, air or air diluted with an inert gas such a nitrogen. Preferably the oxygen-containing gas will be comprised of from about 5 to about 15 percent by volume of molecular oxygen. The duration of the oxygen treatment can be quite brief or can be extended for a period of a few days. Generally, such treatment is for a period of from about 4 hours to 36 hours.

While it is not necessary to employ superatmospheric pressures in the oxygen treatment step, it is usually convenient to do so. Thus, total pressures up to about 300 psig can be employed although normally the total pressure is maintained at a level below 100 psig. Preferably such treatment is carried out at a pressure in the range of about 25 to about 60 psig.

After the oxygen treatment step, the catalyst is purged of molecular oxygen. The purging can be conducted by any of the techniques well known in the art such as, for example, by flowing an inert gas such as nitrogen through the catalyst.

After being purged of molecular oxygen, or simultaneously therewith, the catalyst is cooled to a temperature in the range of about 600°F to about 1,000°F, preferably about 800°F to about 840°F.

After the catalyst has been freed of molecular oxygen and cooled, it is then activated in substantially the same manner as described previously.

As described in detail above, activation is accomplished by simultaneously contacting the catalyst with hydrogen and a chlorine-containing compound, and thereafter purging the catalyst with hydrogen.

The following examples are intended to be illustrative of the present invention and to teach one of ordinary skill in the art to make and use the invention. These examples are not intended to limit the invention in any way.

EXAMPLE I

This example demonstrates activation of a reforming catalyst by reducing the catalyst with hydrogen while contacting the catalyst with a chlorine-containing compound.

A stainless-steel reactor having an inner diameter of about 0.75 inches and a height of about 28 inches was filled with a bottom layer of about 13.5 inches of inert alumina particles having a surface area of 1 m²/g or less, a middle layer of about 6 inches (20 cubic centimeters) of R-56 Pt/alumina reforming catalyst marketed by UOP, Des Plaine, IL (containing about 0.25 weight percent platinum, about 0.40 weight percent rhenium and about

1.0 weight percent Cl on gamma alumina), and a top layer of about 8 inches of inert alumina.

The catalyst system was activated at 940°F by introducing hydrogen at 200 psig while adding perchloroethane at 32 microliters/hr for 15 minutes to give 0.2 weight percent chloride on the catalyst.

Thereafter, a liquid naphtha feed having 23% paraffins, 30% isoparaffins, 8% aromatics and 39% naphthenes was introduced into the reactor at a liquid-volume hourly space velocity of 2.0 hr^{-1} . The reaction pressure was about 200 psig. The reaction temperature was about 860°F. The liquid naphtha had an initial boiling point of 177°F and an end point of 258°F and an average molecular weight of 99.8.

The liquid naphtha feed was added in an amount such that the hydrogen to hydrocarbon ratio was 4.0.

Perchloroethane was then added to this system in an amount of 1.3 ppm to the hydrocarbon feed.

The initial feed had a C5+ content of 100%, a RON value of 62 and a relative octane number of 100%.

After 24 hours of operation of the system, which did not have a hydrogen purge treatment prior to the introduction of the feed, the results were as follows:

C5+ product	RON	Relative OCT-Bbl, %	By-Product ⁽¹⁾
74.6	84.4	102	1.96
⁽¹⁾ methane plus propane mol % in off gas			

EXAMPLE II

This example demonstrates activation of a catalyst in the manner shown in Example I, except after the hydrogen reduction and chloride treatment the catalyst was purged with hydrogen. The hydrogen purge, at a rate of 1.3 SCF per hour, was carried out for a period of 2.0 to 7 hours at about 940°F and 200 psig following the activation of the catalyst system with perchloroethylene addition and prior to the introduction of the naphtha feed to the system.

The results of the test were as follows:

Purge Time (hrs)	C5+ product	RON	Relative OCT-Bbl, %	By-Product ⁽¹⁾
2	84.4	86.2	117	1.47
2.5	86.9	84.5	118	1.63
7	59.9	86.1	83	2.36
⁽¹⁾ methane plus propane mol % in off gas				

The above data demonstrates that the use of a hydrogen purge at a rate of 1.3 SCF of hydrogen per hour for a period up to about 2.5 hours (a time sufficient to provide 4,600 cubic feet of hydrogen per cubic foot of catalyst) after activation of the catalyst system by the addition of chlorine and prior to the introduction of a naphtha feed results in a reforming system which has reduced cracking of product while achieving an increase in RON values. The data further demonstrates that purge times which expose the catalyst to either too much or too little hydrogen produce less favorable result.

EXAMPLE III

This example demonstrates activation of a catalyst in the manner shown in Example I, except the reaction temperature is 840°F rather than 860°F.

After 24 hours of operation of the system, which did not have a hydrogen purge treatment prior to the introduction of the feed, the results were as follows:

C5+ product	RON	Relative OCT-Bbl, %	By-Product ⁽¹⁾
86.6	80.4	112	1.55
⁽¹⁾ methane plus propane mol % in off gas			

EXAMPLE IV

This example demonstrates activation of a catalyst in the manner shown in Example III, except after the hydrogen reduction and chloride treatment the catalyst was purged with hydrogen. The hydrogen purge, at a rate of 1.3 SCF per hour, was carried out for a period of 0.5 to 22 hours at about 940°F and 200 psig following the activation of the catalyst system with perchloroethylene addition and prior to the introduction of the naphtha feed to the system.

The results of the test were as follows:

Purge Time (hrs)	C5+ product	RON	Relative OCT-Bbl, %	By-Product ⁽¹⁾
0.5	91.7	83.9	124	0.92
22	61.6	68	68	2.00
⁽¹⁾ methane plus propane mol % in off gas				

The above data demonstrates that the use of a hydrogen purge at a rate of 1.3 SCF of hydrogen per hour for a period of 0.5 hours (a time sufficient to provide 920 cubic feet of hydrogen per cubic foot of catalyst) after activation of the catalyst system by the addition of chlorine and prior to the introduction of a naphtha feed results in a reforming system which has reduced cracking while achieving an increase in RON values. The data further illustrates that purge times which expose the catalyst to either too much or too little hydrogen, produce less favorable results.

The specific examples herein disclosed are to be considered as being primarily illustrative. Various changes beyond those described will no doubt occur to those skilled in the art; and such changes are to be understood as forming a part of this invention insofar as they fall within the spirit and scope of the appended claims.